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On Discerning Intermolecular and Intramolecular Vibrations in Experimental Acene Spectra

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S Supporting Information

ABSTRACT: The far-infrared spectra of the aromatic hydrocarbons tetracene and pentacene have been calculated using a perturbational-variational method coupled with potential truncation. This shortening of the potential enabled accurate vibration calculations for these large molecules and their respective dimers. Thus, it was possible to identify all IR bands obtained experimentally in the far-IR range by Michaelian et al., as well as to differentiate bands resulting from intermolecular and intramolecular modes of vibration separately, and combined intermolecular + intramolecular vibration modes. Far-IR spectra for smaller acene family members, naphthalene and anthracene, were also computed, and trends in intermolecular vibrations, for the acene family as a whole, were identified. The results obtained illustrate the quality and the detail of the insights realized by interrogating experimental spectra using high-precision, unscaled quantum mechanics computational approaches, and provide a benchmark for future work targeting identification of dominant molecular motifs and intermolecular association phenomena arising in ill-defined hydrocarbons including asphaltenes based on IR and Raman spectral decomposition.

INTRODUCTION

The identification of compounds, compound classes, and principal submolecular motifs present and molecular association phenomena arising in petroleum fractions has been the subject of intense research for many years.^{1–13} However, fluid characterization remains a much debated and critically important topic. Detailed information on the chemical composition and association behavior, particularly of petroleum fractions comprising larger molecules, is needed to predict thermophysical properties, to develop and optimize techniques for hydrocarbon production and refining, and thus to provide opportunities for considerable capital and operating cost reductions and energy efficiency.^{14,15}

Vibrational spectroscopy, a widely used experimental technique for the identification and characterization of molecules in complex chemical environments, has been applied successfully in fields as diverse as biochemistry,^{16–18} agribusiness,^{19–21} interstellar chemistry,^{22,23} as well as the chemistry of materials.^{24–26} In these successful applications, knowledge of the molecules constituting the medium of interest is available. For complex and ill-defined mixtures such as petroleum, a combination of experimental measurement and predictive modeling is required. With the joint progress of modeling techniques^{27–43} and computer technology, quantum mechanical calculations have become an essential and powerful contributor for the identification and discrimination of diverse molecular mixtures.^{44–47} Available mathematical techniques have been tested widely and implemented to serve experimentalists. For example, a recent computational study⁴⁸ focused on petroleum sector applications showed that aromatic and naphthenic motifs, present in large molecules, could in principle be discriminated on the basis of the joint decomposition of Raman and IR spectra using a library of

small aromatic and naphthenic compounds comprising the motifs, whereas proton and ¹³C-NMR provided little information at this length scale.

The work presented here falls within the framework of these mathematical developments to serve identification of complex chemical systems, where the vibrational Schrödinger equation is solved using a double electrical and mechanical anharmonic approximation. Simpler vibration mode calculations, at the harmonic level, are commonly used to predict vibration modes of molecules.^{49–51} However, the resulting error must then be “corrected” using “adapted” multiplicative correction factors that depend on the level of theory employed to solve the Schrödinger equation. These approximate calculations do not allow assignment of modes other than fundamentals, because coupling is not taken into account.

The strategy in the present work is to perform high-precision quantum mechanical calculations based on anharmonic electrical and mechanical approximations to compute spectra for individual molecules and dimers. With this approach, unscaled computed outcomes including harmonics, overtones, combinations, intermolecular vibrations, and other details of spectra can be compared directly with experimental data. The detailed identification of bands in experimental vibration spectra for crude oils and oil fractions may eventually permit the identification of the presence of members of a family of molecules, and the determination of the significance of specific types of molecular interaction arising in these fluids through spectral decomposition. Potential impacts include improved speciation, and hence fluid characterization, and improved

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understanding of association phenomena, such as π - π interactions, and hence transport property prediction.

Acenes, a family of linear polynuclear aromatic compounds possessing only π - π interactions, for which high precision experimental data are available up to pentacene,⁵² comprise a benchmark example that illustrates both the potential and the computational challenges found in this line of inquiry. In this respect, various calculation hypotheses used for these large molecules are presented, together with the impact of these calculation hypotheses on the quality vis-à-vis experimental data and the computational cost of vibrational spectra.

Spectral assignments, determined on the basis of both wavenumber and relative intensity values, are made primarily using computed spectra for individual molecules, and dimers. In addition, on the basis of experimental observation assessing the presence of deformed molecules in tetracene and pentacene crystals, the impact on vibrational spectra of these deformations is evaluated. For some cases, assignment ambiguity is only resolved by tracing trends for vibration modes for the acene family as a whole. In the exposition that follows, each of these topics is tackled sequentially. Calculations with tetracene and pentacene monomers (individual molecules) are treated first. Then, calculations for dimers and the acene family collectively are treated, and the results synthesized.

■ CALCULATION OF VIBRATIONAL SPECTRA OF TETRACENE AND PENTACENE MONOMERS: COMPARISON WITH EXPERIMENTAL DATA

a. Experimental Data. Experimental solid-state infrared spectra of tetracene and pentacene⁵² were acquired at room pressure and temperature. The measurements were performed on as-received samples. Thus, the crystallinity as well as the possible presence of defects in these solids are unknown. For tetracene, two crystal structures have been detected by X-ray diffraction experiments.⁵³ At room temperature and pressure, tetracene is triclinic⁵⁴ with space group $P\bar{1}$. The unit cell contains two independent molecules at the (0, 0, 0) and (1/2, 1/2, 0) inversion sites. Pentacene possesses two polymorphs.^{55,56} One comprises a triclinic layered structure with a herringbone arrangement in the layers, with two equivalent molecules per unit cell.⁵⁷ The other, commonly found at room temperature, adopts a 14.1 Å $d(001)$ -spacing morphology with an inversion center on both molecules in the unit cell.^{58–62}

b. Computational Details: Construction of the Potential Energy Surface. The theoretical process generally developed to resolve the vibrational Schrödinger equation requires two restricted steps: the construction of the potential energy surface (PES) and the resolution of the vibrational equation in order to obtain the harmonic and anharmonic energy levels and consequently the fundamental, overtone, and harmonic wavenumbers characterizing the vibrational spectra. Calculations of the intensities complete the analysis of the spectra.

The potential energy surface is normally constructed by expressing the potential function as a Taylor series expansion using curvilinear displacement coordinates. This series is often truncated at the fourth order. Quadratic, cubic, and quartic force constants are generally obtained by fitting the electronic energy data calculated by *ab initio* methods for various nuclear configurations close to the optimized geometry, or by a finite difference procedure for first or second derivatives of the electronic energy with respect to the nuclear coordinates.⁶³

The Becke three-parameter hybrid functional (B3LYP), a combination of the Becke three-parameter exchange functional (B3)⁶⁴ that includes a mixture of Hartree–Fock exchange with the density functional theory (DFT) exchange–correlation and the LYP correlation functional,⁶⁵ has proven to be a reasonable choice for predicting geometries of aromatic molecules.^{66,67} Stephens et al.⁶⁸ showed that the B3LYP force field yields infrared spectra in very good agreement with experimental data. It has become a standard method for studying vibrational spectra of organic molecules (in the far- and mid-infrared) in the gas phase.^{49,69} Although the B3LYP method coupled with a sufficiently large basis set is a good candidate to study vibrational spectra of polynuclear aromatic molecules,⁵² this standard DFT calculation approach fails to describe long-distance interactions. Consequently, it does not reproduce spectra of interacting molecules^{64,68,70,71} where the interpretation of experimental vibration spectra requires consideration of both isolated and interacting molecules.

In parallel to the development to *ab initio* methods, some progress has been made on implementing the London dispersion force effect into DFT methods, permitting calculations with larger individual molecules and binary pairs. Although the CCSD(T) method in combination with a sufficiently flexible basis set is the most accurate method for reproducing rotation-vibration spectra and to describe intermolecular complexes between aromatic hydrocarbons,⁷² such calculations are impractical for the description of large molecules.⁷³ The MP2 method,⁷⁴ also able to handle dispersive interactions such as π - π stacking,^{75,76} is a computationally intensive and inadequate method for large molecules too.⁷⁷ Chai et al.⁷⁸ developed a functional based on optimized long-range corrected hybrid density functionals,⁷⁹ which employs 100% Hartree–Fock (HF) exchange for long-range electron–electron interactions, ω B97X, to which they added an empirical dispersive interaction correction.⁷⁷ This method was tested by Salzner et al.⁸⁰ for π -conjugated oligomers, and it overcame the difficulties encountered with standard DFT functionals when dealing with interacting π -systems.

In this work, extended basis sets consisting of atomic orbitals expressed as fixed linear combinations of Gaussian functions are employed for the calculations. A split valence 6-311G and 6-311G(p,d)^{81,82} was needed to provide correct calculation outcomes for naphthalene monomer.⁵⁰ Validation tests for the ω B97X-D/6-311G calculation method were also performed. The accuracy of DFT calculations in predicting molecular geometries and vibrational frequencies depends on the density functional employed. Thus, geometrical optimizations were performed for isolated tetracene, using the ω B97X-D functional as well as the more conventional B3LYP method. The computational results are benchmarked using experimental data obtained by Campbell et al.⁵⁴

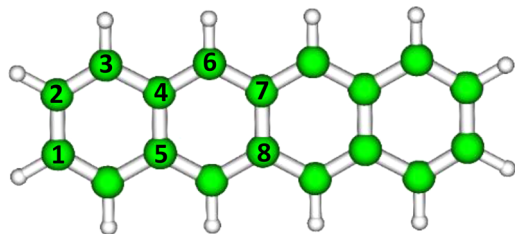
c. Geometrical Parameters of Isolated Tetracene Molecule Optimized Using B3LYP and ω B97X-D Methods Associated with the 6-311G and 6-311G Basis Sets.** Geometry optimizations were performed using two different DFT methods (B3LYP and ω B97X-D) implemented in Gaussian 09 software⁸³ and two basis sets (6-311G and 6-311G**). While Saeki et al.⁵⁰ reported a change of symmetry for the naphthalene molecule, from D_{2h} to C_{2h} (based on MP2/6-311G calculations), optimized geometries for the tetracene molecule, which also possesses D_{2h} symmetry, remained planar with D_{2h} symmetry in all cases. Calculated bond lengths based on different levels of theory are listed in Table 1 along with

Table 1. Method and Basis-Set Dependence of C–C Bond Length (Å)^a for Tetracene Monomer and Mean Deviation from Experimental Data^b

	C1–C2	C2–C3	C3–C4	C4–C5	C4–C6	C6–C7	C7–C8	mean deviation
experimental data ^a	1.46	1.38	1.42	1.42	1.39	1.40	1.46	
B3LYP/6-311G	1.43 (−0.03)	1.36 (−0.02)	1.43 (+0.01)	1.43 (+0.01)	1.40 (+0.01)	1.42 (+0.02)	1.43 (−0.03)	0.02
B3LYP/6-311G**	1.43 (−0.03)	1.36 (−0.02)	1.43 (+0.01)	1.43 (+0.01)	1.40 (+0.01)	1.42 (+0.02)	1.43 (−0.03)	0.03
ω B97X-D/6-311G	1.43 (−0.03)	1.37 (−0.02)	1.44 (+0.02)	1.45 (+0.03)	1.39 (=)	1.41 (+0.01)	1.45 (−0.01)	0.02
ω B97X-D/6-311G**	1.43 (−0.03)	1.37 (−0.02)	1.44 (+0.02)	1.45 (+0.02)	1.39 (=)	1.41 (+0.01)	1.45 (−0.01)	0.02

^aThe values in parentheses indicate the difference between the calculated and experimental bond lengths. ^bExperimental bond lengths are obtained from ref 54.

their mean deviations from experimental measurements (see also Figure S1 in the Supporting Information). For these calculations, irrespective of the level of theory or the basis set employed, the agreement between computed bond lengths and experimental values is very good. The mean deviation from experimental values is less than 0.03 Å. These results show clearly that the choice of calculation method and basis set has a minor influence on the geometry of tetracene. Results obtained with pentacene monomer are comparable. This basis-set insensitivity for DFT approaches has been illustrated previously.⁸⁴ It permits the use of smaller basis sets and provides an additional argument for the use of DFT over *ab initio* methods, since larger molecules can be treated. The ω B97X-D method, which contains a London-force dispersive term and is reliable for acene molecules, was selected for this study, and it is used in combination with the 6-311G basis set.

**Figure 1.** Representation of a tetracene molecule.

d. Truncation of the Potential Energy Surface.

Vibrational spectroscopy modeling outcomes, whether at the harmonic or anharmonic level, are directly dependent on the

quality of the potential energy surface (PES). The PES is generally expressed as a polynomial determined by adjustment of a set of geometric structures of a molecule. Consequently, this type of calculation depends on the quality of the electronic calculations of the wave functions as well.

While it is now common to determine the PES accurately for small molecules (3–5 atoms),⁸⁵ this becomes almost infeasible when solving anharmonic approximations for systems of larger size where powerful computers are required to implement rigorous methods for solving the Schrödinger equation. Truncation of the analytical shape of the PES (degree of the potential) is unavoidable. For example, a tetracene molecule contains 30 atoms, and therefore has $3N - 6 = 84$ normal modes apart from the translational and rotational modes. 84 modes with an order 4 necessitate the calculation for 22 171 force constants. Thus, truncation of the PES constitutes the first part of this study. Four levels of approximation were tested. The fundamental bands obtained in these cases, using second-order perturbation theory (PT2), are compared to the ones calculated with the complete PES. The conclusions of these tests are summarized in Table 2 and elaborated here:

(i) It is possible to truncate the PES in terms of only significant coupling. For this, it suffices to identify the constant of cubic and quartic force of very low intensity.

(ii) We have realized calculations by successively suppressing force constants smaller than 2, 5, 8, and 10 cm^{-1} . This study showed that the discrepancy regarding the calculated frequencies of fundamental modes is not higher than 2.8 and 3.3 cm^{-1} for the combination modes. These disparities are small enough to justify use of this approximation. Moreover, this allows setting 7, 33, 48, and 55% of the forces to zero when

Table 2. Force Constant Truncation and Separation of Low and High Mode Influences on the Values of Wavenumbers of Tetracene Monomer Based on a Force Field Extracted from Calculations at the ω B97X-D/6-311G Level of Theory

polynomial degree	2	4 ^c	4 ^c	4 ^c	4 ^c	6 ^c
force constants (<i>f</i>) considered		all <i>f</i> considered	$f > 5 \text{ cm}^{-1}$	$f > 5 \text{ cm}^{-1}$	$f > 5 \text{ cm}^{-1}$	all <i>f</i> considered
number of modes	84	84	84	20	10	10
mode (symmetry)	exp. values ^a					
ν_1 (B_{3u})	57	48	53	62	60	69
ν_2 (A_u)	93	90	96	107	103	107
ν_3 (B_{1g})	154	150	154	161	160	164
ν_4 (B_{1u})	166 ^b	168	160	169	169	171
ν_5 (B_{2g})	196	194	200	204	202	203
ν_6 (B_{3u})	271 ^b	276	272	277	284	284
ν_7 (B_{3g})	316	317	305	317	317	319
ν_8 (A_g)	322 ^b	324	322	323	322	322
ν_9 (A_u)	327	325	332	330	328	331
ν_{10} (B_{1g})	389	382	386	395	392	394

^aExperimental values of wavenumbers are obtained from ref 52. ^bExperimental uncertainty equal to $\pm 2 \text{ cm}^{-1}$. ^cWavenumbers calculated using second-order perturbation theory (PT2).

suppressing constants smaller than 2, 5, 8, and 10 cm^{-1} , respectively, thus saving computational cost and enabling variational calculations for larger systems. Note, however, that this restriction is potentially acceptable only without any Fermi or Darling–Dennison resonance, often characterized totally or partially using these finest couplings. Regarding the time saved when suppressing the forces lower than 10 cm^{-1} while the wavenumber values of fundamental modes remain the same, it would be tempting to truncate the PES this way, but this would represent a serious risk of losing information regarding combination bands. Indeed, our tests show that the upper limit of the truncation to maintain all the information on combination bands consists of considering all the forces higher than 5 cm^{-1} . Removal of these forces would disregard combinations whose weight is less important, resulting in the loss of details that provide a complete interpretation of infrared spectra. Thus, for the following tests of truncature and for calculations presented in this study, we will consider a PES calculated with force constants higher than 5 cm^{-1} .

(iii) The second truncation tested in this work consists of separating modes of high and low wavenumber. We assumed that the low-wavenumber modes (mainly torsion modes) were weakly coupled with bending and scissoring modes occurring at higher wavenumber. This hypothesis has been tested and validated on the basis of calculations on the tetracene molecule (Table 2); indeed, each of the first 15 modes calculated using the PES truncated to torsional and libration modes are in perfect adequacy with values calculated *a posteriori* using complete PES. Figure S2 in the Supporting Information shows that the effect of suppressing the higher modes has a minor incidence on the values of the fundamental modes calculated at the vibrational MP2 level. However, suppressing modes higher than 500 cm^{-1} was even beneficial when considering combination bands. Indeed, calculations can cause the modes of higher frequencies, with large amplitude, to interfere artificially on these “soft modes” localized in the range [0, 400] cm^{-1} , giving rise to nonphysical frequencies. Another consequence of the artificial interactions between these modes is that combination modes of low wavenumber are embedded in the parasitic values of frequencies, and then result in loss of information.

(iv) We have also tested the association of a complete grid of points of the surface at the fourth order, with a fit at the sixth order on the diagonal, in order to refine coupling between closely coupled modes. Because this refining is demanding in terms of computing time, we applied it on the 10 first modes only, supposing that the modes of higher wavenumber were correctly described with a polynome of fourth order. The results, listed in Table 2, show that this refining does not bring any improvement for describing the vibrational spectrum of tetracene.

The last approximation used in this work consists of applying a second adjustment of the PES; this adjustment would aim to eliminate the weight of vibrators we excluded in point i, to refine the fit of remaining force constants so that they reflect, *in fine*, only the couplings between remaining modes.

If this had proved to be an acceptable approximation, we could have extended this methodology to the study of specific molecule vibrators of undifferentiated dimension, by calculating a limited number of specific points to sought vibrators.

In accordance with what Bowman et al.⁸⁶ and Begue et al.⁸⁷ showed on the study of water clusters, our calculations show

that only the separation between libration modes with all other active modes is possible.

To conclude, we have shown that partial truncation of PES could be effective for our study. This truncation consists of extracting from a global fit the force constants of higher order that allow the description of the problem beyond the harmonic approximation. These constants can then be directly used to solve Schrödinger’s vibrational equation.

This methodology was used in the remainder of our work for the characterization and assignment of the spectra of tetracene and pentacene molecules.

e. Resolution of the Vibrational Schrödinger Equation (Noncore Bands with Intensity). The presence of overtones and combination bands in infrared spectra is a manifestation of the breakdown of the double-harmonic approximation, usually implemented in electronic structure codes. The calculation of intensities is essential in order to determine which non-fundamental transitions are active and to assign all experimental bands. Since both mechanical (anharmonicity of the potential) and electrical (nonlinear dependence of the dipole moment on the normal coordinates) anharmonicities are expected to give nonzero intensities to nonfundamental transitions, these two effects must be considered in the treatment of transition energies and related vibrational wave functions. The wavenumber calculations in the mechanical anharmonic approximation were carried out using a variational method developed by Bégué et al.⁸⁵ Using this method, implemented in the P_Anhar.v2.0 program,⁸⁸ it was possible to compute all the vibrational frequencies (fundamental, combination bands, and overtones) that contribute to the mid- and near-infrared spectrum of tetracene and pentacene. In addition, the activity of each mode was also calculated in the electrical anharmonic approximation using a method developed by Baraille et al.,^{87,89} which is also available in the same software.

The experimental spectra of tetracene and pentacene, extracted from the work of Michaelian et al.,⁵² are shown in Figure 2. The corresponding calculated infrared-active transitions—both fundamental and combination—overlay the data and are listed in Tables 3 and 4 where they are further compared with available experimental and theoretical values. A

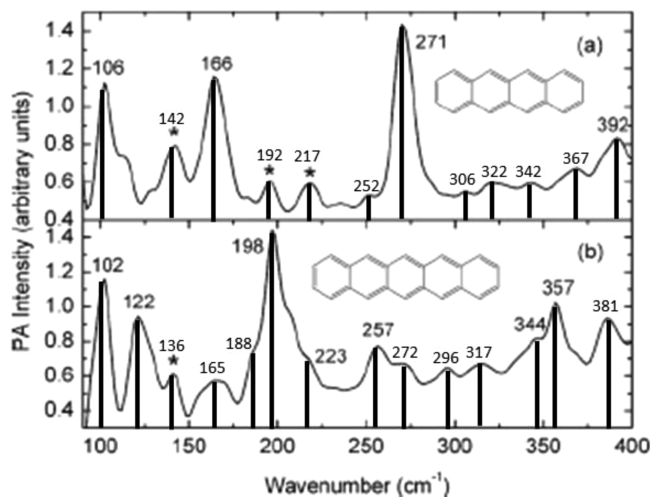


Figure 2. Far-infrared PA spectra of (a) tetracene and (b) pentacene. Spectra were acquired using phase modulation at a frequency of 5 Hz and an amplitude of 15λ ($\lambda = 0.6328 \mu\text{m}$), extracted from Michaelian et al.⁵²

Table 3. Experimental and Calculated Wavenumbers (cm^{-1}) for Fundamental and Combination Vibrational Modes of Tetracene^d

Mode	Type of vibration	Symmetry	Experimental values	Predicted bands					
			Michaelian <i>et al.</i> (52)	Michaelian <i>et al.</i> (52)	Cataldo <i>et al.</i> (90)	Harmonic (cm^{-1})	Anharmonic (PT2) (cm^{-1})	Theoretical IR Intensities ^a (km/mol)	
			106(m)						
v3	Wagging	B _{1g}	142(m)				154R ^c	153R ^c	0.01 ^b
3v1		B _{3u}					171	132	0.003
v4	Scissoring	B _{1u}	166(s)	B _{1u}	166	160	168	164	1.3
v5	Wagging	B _{2g}	197(w)				196R ^c	199R ^c	0.05 ^b
v1+v3		B _{2u}					211	201	0.005
v1+2v2		B _{3u}	217(w)				243	228	0.001
v2+v3		B _{1u}	252(vw)				247	247	0.002
v6	Wagging	B _{3u}	271(m)	B _{3u}	278	266	276	275	1.03
v2+v5		B _{2u}	306(vw)				289	293	0.002
v8	Stretching	A _g	322(w)				324R ^c	322R ^c	0.002 ^b
v1+2v3		B _{3u}	342(w)				365	353	0.003
v1+v8		B _{3u}	367(m)				381	379	0.003
v10	Alternate wagging	B _{1g}	392(m)	B _{1g}	388R	374R	389R	385R	0.04 ^b
2v1+v6		B _{3u}					390	380	0.004

^aTheoretical intensities obtained using perturbational-variational calculations, with truncated potential on the 20 first modes. ^bTheoretical intensities obtained for the distorted molecule, at the harmonic level. ^cThe letter R indicates Raman-active modes. ^dColored lines indicate cases with double-hypotheses.

Table 4. Experimental and Calculated Wavenumbers (cm^{-1}) for Fundamental and Combination Vibrational Modes for Pentacene^d

Mode	Type of vibration	Symmetry	Experimental values	Predicted bands					
			Michaelian <i>et al.</i> (52)	Michaelian <i>et al.</i> (52)	Cataldo <i>et al.</i> (90)	Harmonic (cm^{-1})	Anharmonic (PT2) (cm^{-1})	Theoretical IR Intensities ^a (km/mol)	
v3	Wagging	B _{1g}	102 m	B _{1g}	105R ^c	100R ^c	104R ^c	97R ^c	0.003 ^b
v4	Scissoring	B _{1u}	122 m	B _{1u}	121	117	123	122	1.0
v1+v3		B _{2u}	141 m				142	144	0.004
v2+v3		B _{1u}	165 vw				178	164	0.0005
v6	Wagging	B _{3u}	198 m	B _{3u}	198	189	198	194	1.4
v3+v4		A _u	223 vw	A _u			224		0
v2+v5		B _{2u}					226	222	0.002
v9	Symmetric stretching	A _g	257 w	A _g	266R ^c	258R ^c	269R ^c	265R ^c	0.03 ^b
v1+2v3		B _{3u}					247	251	0.0006
2v1+v6		B _{3u}	272 vw				274	274	0.0007
v3+v6		B _{2u}	296 w				302	291	0.003
v1+v9		B _{3u}					307	292	0.003
v2+v7		B _{3u}	317 w				321	316	0.0009
v3+v8		B _{1u}	344 vw				351	346	0.03
v12	Scissoring	B _{1u}	357 w	B _{1u}	369	357	375	374	0.1
v13	Alternate wagging	B _{3u}	381 m				385	374	0.03

^aTheoretical intensities obtained using perturbational-variational calculations, with truncated potential on the 20 first modes. ^bTheoretical intensities obtained for the distorted molecule, at the harmonic level. ^cThe letter R indicates Raman-active modes. ^dColored lines indicate cases with double-hypotheses.

threshold intensity for IR activity of 0.001 km/mol was chosen for the calculations. This threshold is justified following a discussion of band assignments for each molecule, and presentation of vibrational bands common for the two molecules.

For tetracene molecules, three bands were identified previously in the range 100–400 cm^{-1} .^{52,91} Two bands (166 and 271 cm^{-1}) were noted as infrared-active, and one (388 cm^{-1}) was assigned as Raman-active (Table 3). By calculating

intensities for all bands, including combination bands and overtones, eight additional infrared-active vibrations (seven combinations and one second-harmonic) are identified, and the band at 388 cm^{-1} previously assigned as Raman-active is identified as an infrared-active band arising from the B_{3u} mode plus an overtone. Three similar cases where a vibrational transition could correspond either to a combination band, an overtone, or a Raman-active band are also shown in rows with colored backgrounds in Table 3. The hypothesis of Raman

activity of these bands requires further elaboration prior to performing a definitive assignment of these bands, based on their intensities, and where necessary wavenumber trends for bands common to both tetracene and pentacene.

Raman-Transition Activity in Infrared Spectra: Molecular Deformation. Raman transitions can become active in infrared spectra, when a molecule is deformed in the solid state. For example, Homes et al.⁵⁸ showed experimentally that acene molecules in the solid state are not planar and attributed this observation to crystal-packing forces that induce deformation. Further, deformation affects measured lattice-phonon frequencies, permitting discrimination of different solid phases of tetracene⁵³ and pentacene.^{55,92} In order to underpin this contention, calculations on the deformed tetracene molecule shown in Figure 3 were performed. In this conformation, the

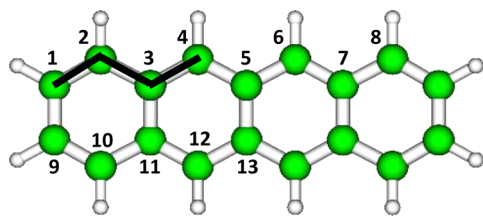


Figure 3. Representation of a tetracene molecule with a deformed dihedral angle $C_1-C_2-C_3-C_4$ shown in black.

molecule does not possess D_{2h} symmetry, but the molecule does still possess symmetry, because the deformation concerns dihedral angles $[C_1-C_2-C_3-C_4]$, $[C_2-C_3-C_4-C_5]$, $[C_4-C_5-C_6-C_7]$, $[C_5-C_6-C_7-C_8]$, $[C_9-C_{10}-C_{11}-C_{12}]$, and $[C_{10}-C_{11}-C_{12}-C_{13}]$, equal to 175, 178, -178, -178, -178, and -178°, respectively, instead of $\pm 180^\circ$ in the case of the molecule in the D_{2h} conformation. The bond lengths are unchanged. Wavenumbers of fundamental modes and associated intensities calculated with the harmonic approximation for deformed tetracene are listed in Table S1 in the Supporting Information. The calculated vibrational spectrum exhibits only positive wavenumbers, which attests that this

deformation is located at a minimum of the potential energy surface. The intensities corresponding to the R-denominated bands of the deformed molecule, calculated at the harmonic level, are given in Table 3. For the deformed molecule, modes that would normally be inactive in infrared (Raman-active or silent modes) present significant intensities, ranging from 0.002 to 0.05 km/mol.

Assignment Based on Band Intensity Value. For tetracene, three transitions for which there is a double-hypothesis (142, 192, and 392 cm^{-1}) can be resolved on the basis of relative intensity values. The band at 142 cm^{-1} could be attributed to a second harmonic whose intensity would be 3 times lower than the Raman-active one that becomes optically active for a deformed molecule. The band at 192 cm^{-1} could be a combination of an IR-active mode and a Raman-active mode, whose intensity would be 10 times lower than a Raman-active band calculated at 196 cm^{-1} . The band at 392 cm^{-1} could either be a combination of an overtone and an infrared-active mode or a Raman-active band 10 times more intense. From the relative intensities, in each case, the hypothesis of a Raman band becoming active is more likely. For other transitions, there is a single hypothesis. The transition, denoted ν_8 , at 322 cm^{-1} , is a Raman-active band, while bands at 252, 306, and 342 cm^{-1} are combination bands.

The interpretation of results for pentacene, Table 4, is similar. Eleven bands arise from combination vibrations. One, found experimentally at 257 cm^{-1} , was previously assigned to a predicted Raman-active mode. Calculations on a deformed pentacene molecule lead to an IR-active band at a comparable wavenumber and with a comparable intensity. Thus, this one assignment remains uncertain on the basis of these calculations.

Use of Wavenumber Trends to Assign Bands. The trends for normal modes of infrared-active vibrations found for both tetracene and pentacene molecules (see types of vibration in Tables 3 and 4) are shown in Figure 4 where they are also compared with experimental data. By viewing the wavenumbers in this way, wavenumber assignments for the two molecules can be evaluated in parallel. From the similarity of the trends for tetracene and pentacene, as well as the proximity of the

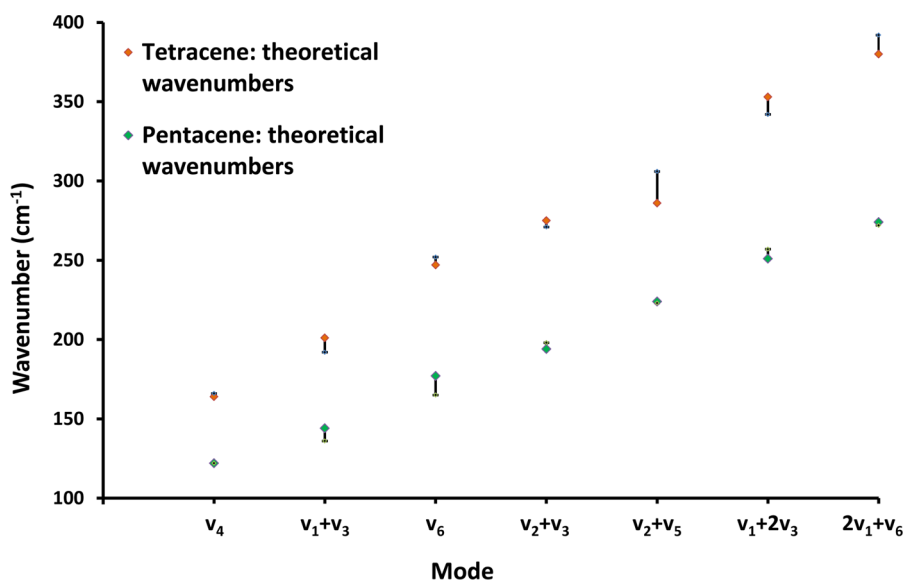


Figure 4. Comparison of theoretical and experimental wavenumbers of tetracene and pentacene for the same normal modes of vibration. The deviations between experimental and theoretical wavenumbers are represented with dark bars.

respective computed wavenumbers to experimental ones, the assignment choices—combination bands over Raman-active ones—are supported and this outcome is independent of intensity values. Only one band remains unassigned in the experimental spectrum of tetracene, 106 cm^{-1} . Michaelian et al.⁵² suggested that it originated from lattice vibration. The pentacene band attributed to a Raman-active vibration, 102 cm^{-1} , may be of the same type as well. These bands and intermolecular bands more broadly are addressed by performing calculations with tetracene and pentacene dimers.

■ CALCULATION OF VIBRATIONAL SPECTRA OF ACENE DIMERS: COMPARISON WITH EXPERIMENTAL DATA

π -Stacking is responsible for the cohesive structure of acene crystals.^{54,60,93–95} Dimers constitute a benchmark for describing π -stacking phenomena. In order to generalize the effect of π -stacking on intermolecular vibrations for the acene family as a whole, the calculations described here include those for naphthalene and anthracene dimers in addition to the tetracene and pentacene dimers. Inclusion of these smaller dimers also permits calibration of the computational methodology employed.

a. Calculated Geometry and Wavenumbers of Dimers and Comparison with Experimental Data. Far-infrared spectra provide information about intermolecular and inter-ring interactions and about π -stacking in particular. Hineno et al.⁹⁶ measured spectra of naphthalene and anthracene in the region $190\text{--}50\text{ cm}^{-1}$ at liquid helium temperature and observed that some bands were due to lattice vibrations. In the same way, bands in the experimental vibrational spectra of tetracene and pentacene that cannot be interpreted on the basis of computed wavenumbers for a single molecule are expected to arise from lattice modes. A first step toward understanding π - π interaction of acene compounds consists of considering the interaction of acene dimers. The choice of the calculation method for describing these electronic interactions is justified, and then, the differentiation of intermolecular and intramolecular vibrations in the acene family is addressed.

Naphthalene Dimer. To assess the performance of the ω B97X-D method for π - π interaction calculations involving aromatic molecules, the robustness of the ω B97X-D/6-311G method for naphthalene dimers, for which a broad literature exists,^{50,97–100} is discussed in light of results based on reliable data and accurate computational results. For example, Gonzalez et al.⁹⁷ and Tsuzuki et al.⁹⁸ calculated ground-state structures and binding energies of van der Waals dimers of naphthalene using an aromatic intermolecular interaction model, at the MP2 and CCSD(T) levels of theory, respectively, using Pople's basis sets of various sizes. Their calculations yield two low energy dimers of similar energy, D_{2d} (crossed) and C_{2h} (parallel-displaced), and two C_{2v} T-shaped structures. Walsh et al.,⁹⁹ using calculations at the MP2/6-31*(0.25) level, found eight stationary points for naphthalene dimers: the four reported previously and four others of lower symmetry. Following these studies, Saeki et al.⁵⁰ performed vibrational analysis of naphthalene dimers at the MP2/6-311G level of theory and their results are compared with the values obtained in this work.

In the present work, geometry optimizations of naphthalene dimers in various configurations—parallel-displaced (PD), parallel-displaced according to C_{2h} symmetry, crossed-parallel according to C_2 and D_{2h} symmetry—using the ω B97X-D method with the 6-311G and 6-311G(p,d) basis sets were

performed. All stationary points were characterized by frequency calculations to ensure that the optimized structures were located at a local minimum. The basis-set superposition error (BSSE) correction developed by Simon et al.¹⁰¹ was included in the calculations for all complexes. Calculations without this correction were carried out to characterize the importance of this error. Negative binding energies indicate exothermic complex formation.¹⁰² Analysis of dimer configurations showed that only the low-level symmetry configurations, C_2 and parallel-displaced (PD), represented in Figure 5a and b, respectively, correspond to minima on the PES.

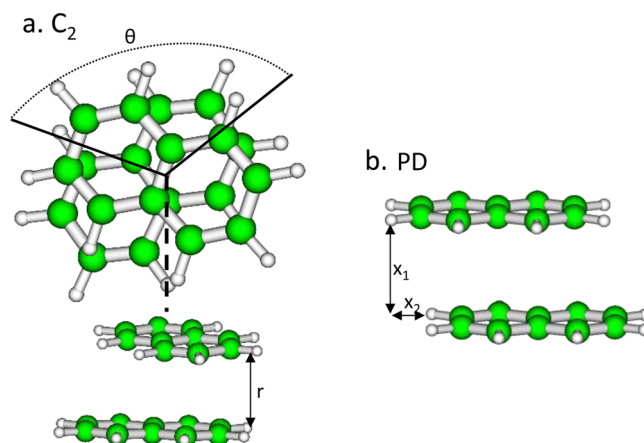


Figure 5. Optimized geometries at the ω B97X-D/6-311G level of theory for naphthalene dimers in the most stable configurations (a) C_2 and (b) parallel-displaced (PD).

Vibrational calculations for the symmetric configurations exhibited negative frequencies. This outcome is in agreement with Walsh et al.⁹⁹ who showed that structures of low symmetry are the most stable. The center of mass separation (r), rotation angle (θ) for the C_2 structure, and distance of relative translation (x_1 and x_2) for the PD structure are given in Table 5. For these two configurations, the interaction energies

Table 5. Geometry Dependence of Basis Sets on Structures PD and C_2 Shown in Figure 5

	C_2		PD		
	θ (deg)	r (Å)	r (Å)	x_1 (Å)	x_2 (Å)
6-311G	129	3.38	3.52	1.12	1.26
+CP ^a	126	3.41	3.44	1.08	1.29
6-311G**	130	3.45	3.32	1.07	1.27
+CP ^a	133	3.43	3.65	1.05	1.30
cc-pVDZ+CP ^a	134	3.45	3.49	1.11	1.28
Saeki et al. ⁵⁰	135	3.46	3.40	1.03	1.31

^aCP refers to the keyword “counterpoise” used in Gaussian to take BSSE correction into account in the calculations.

at the different levels of theory considered, with and without BSSE correction, as well as the first harmonics, are reported in Table 6. Interaction energies and associated wavenumbers were obtained using ω B97X-D and 6-311G, 6-311G**, and cc-pVDZ basis sets. Computed values are compared with results obtained by Walsh et al.,⁹⁹ Saeki et al.,⁵⁰ and Rubes et al.¹⁰⁰ BSSE correction lowers the absolute value of the energy by 1.5 kcal.mol^{-1} on average, and represents only 0.04% of the binding energy. Thus, BSSE correction does not affect the

Table 6. Interaction Energies in kcal/mol and Associated Wavenumbers (cm^{-1}) of Optimized Naphthalene Dimers

naphthalene		C_2		PD	
ω B97X-D/6-311G (this work)	without counterpoise	-9.00	14	-8.72	30
	with counterpoise	-7.55	8	-7.28	33
ω B97X-D/6-311G** (this work)	without counterpoise	-9.45	16	-9.16	32
	with counterpoise	-7.86	12	-7.58	29
ω B97X-D/cc-pVDZ (this work)	with counterpoise	-7.52	12	-7.22	25
MP2/6-31G* (0.25) (Walsh et al. ⁹⁹)	with counterpoise	-8.15		-7.68	
MP2/cc-pVDZ (Saeki et al. ⁵⁰)	without counterpoise	-5.30	22	-5.00	8
	with counterpoise	-6.24		-5.89	
DFT/CCSD(T) (Rubes et al. ¹⁰⁰)		-6.23		-5.96	
SAPT0/aug-cc-pVDZ (Hohenstein et al. ¹⁰³)				-7.27	

order of stability of the tested configurations. This outcome is in full agreement with the cited prior work. Further, there are no trends for the impact of basis set and BSSE correction on the geometry of the dimer configurations. However, by calculating the mean difference of computed values, this work, with those of Saeki et al.,⁵⁰ shows the 6-311G basis set combined with BSSE correction to be a preferred option.

Vibration Modes of Naphthalene Dimers. Wavenumber calculations for stable dimers were performed using the ω B97X-D/6-311G method with BSSE correction. From vibrational analysis of the monomer (Table S2 in the Supporting Information) and the dimer (Table 7), the vibrational modes arising from intra- and intermolecular modes can be discriminated. The assignments are reported in Table 7 for sets of calculations performed as part of this work

and from Saeki et al.⁵⁰ Calculated values are compared with experimental data when available. The calculations were performed with counterpoise correction for both the C_2 and PD configurations. The calculated intra- and intermolecular wavenumbers, irrespective of calculation details, are readily distinguished for naphthalene. The intermolecular wavenumbers are at $\sim 100 \text{ cm}^{-1}$ and lower, while the intramolecular wavenumbers are greater than $\sim 100 \text{ cm}^{-1}$.

b. Pi-Stacking in the Acene Family. Optimization at the ω B97XD/6-311G level with BSSE correction gives rise to the expected parallel tilted configuration¹⁰⁵ for anthracene, tetracene, and pentacene dimers, as shown in Figure 6. This configuration closely resembles that found in triclinic tetracene⁵⁴ and pentacene⁵⁸ crystals. The associated interaction energies are reported Table 8. Other configurations were found to be unstable. They possess negative calculated harmonic values and were clearly not located at a minimum of the PES. Thus, only naphthalene has a more energetically stable C_2 configuration. Computations related to both the C_2 and PD configurations for naphthalene are shown in Table 7. The PD configuration results are used for comparisons with other acene family members where it is the most stable configuration.

For a given geometry, in this case parallel-tilted, the binding interaction energy increases with the number of interacting cycles in a nonlinear manner. This result is consistent with previous studies by Hohenstein et al.¹⁰³ and Grimme et al.¹⁰⁶ who compared interaction energies of naphthenic and aromatic dimers of increasing size using the B2PLYP/TZV(2d,p) level of theory. They found that arene dimers were strongly preferred over equivalent naphthenic dimers, especially for anthracene and tetracene which were stabilized by 3–4 kcal·mol⁻¹. The calculations reported here are in accordance with these prior results and confirm the prevalence of π -stacking interactions.

The wavenumbers assigned to intermolecular (Table 9) and intramolecular (Table 10) modes of vibration for the acene family were calculated at the ω B97XD/6-311G level of theory using BSSE correction. The characteristic wavenumbers for

Table 7. Calculated Wavenumber (cm^{-1}) of the Intra- and Intermolecular Vibrational Modes of the Naphthalene Dimer in Configurations a and b at the ω B97X-D/6311G Level with BSSE Correction, Compared with Frequencies Obtained at the MP2/cc-pVDZ Level in Saeki et al.⁵⁰

configuration C_2 ω B97X-D/6-311G ^a		Saeki et al. ⁵⁰ MP2/cc-pVDZ ^b	configuration PD ω B97XD/6-311G ^a		Saeki et al. ⁵⁰ MP2/cc-pVDZ ^b	experimental values ¹⁰⁴
harmonic	PT2		harmonic	PT2		
Intramolecular						
184	181	185	188	181	180	176
192	187	188	196	188	187	195
199	198	203	207	201	201	
203	202	205	208	201	207	
378	370	351	380	372	351	
378	371	351	381	372	351	359
404	398	386	405	396	385	
405	398	387	405	397	387	
Intermolecular						
14	16	22	33	29	8	
32	34	45	36	35	39	
50	50	64	55	51	54	
78	83	95	75	68	81	
85	89	103	88	85	100	
94	99	106	101	98	110	

^aWith counterpoise correction. ^bWithout counterpoise correction.

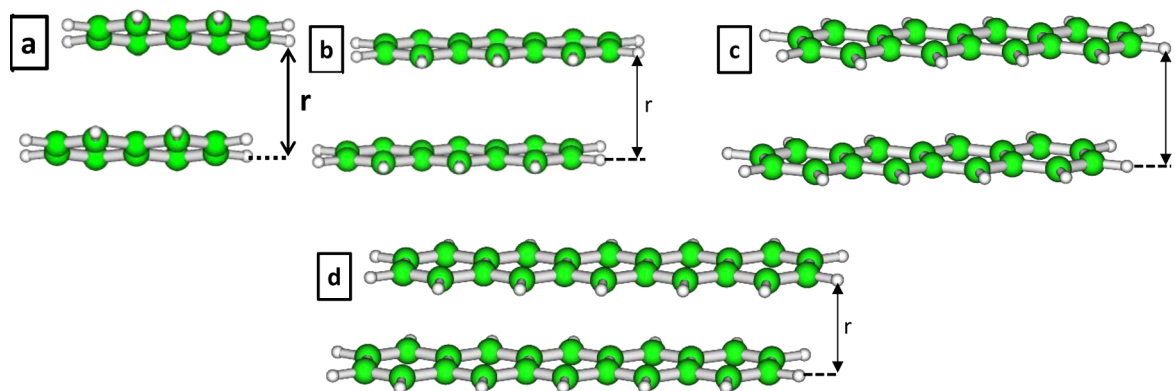


Figure 6. Acene dimers in the parallel-displaced (PD) configuration: (a) naphthalene; (b) anthracene; (c) tetracene; (d) pentacene.

Table 8. Intermolecular Distance and Interaction Energies Calculated for Acene Dimers

dimers	naphthalene	anthracene	tetracene	pentacene	
intermolecular distance r (Å)	3.41	3.27	2.96	2.98	
interaction energies (kcal/mol)	ω B97XD/6-311G + BSSE correction	−7.53	−11.52	−15.73	−20.44
	Hohenstein et al. ¹⁰¹ SAPT0/aug-cc-pVDZ	−7.27	−12.24	−17.49	−22.91

Table 9. Calculated Intermolecular Vibrational Modes (cm^{-1}) and Their Intensities for Naphthalene, Anthracene, Tetracene, and Pentacene Dimers

type ^a	naphthalene	anthracene	tetracene	pentacene
rotation 1/2	33(0.04)	22(0.004)	15(0.002)	19(0.002)
translation 1/2	36(0)	34(0)	31(0)	27(0)
total rotation + translation 1/2	55(0)	57(0)	52(0)	57(0)
	75(0.21)	75(0.05)	87(0.14)	74(0.07)
rotation 1/2	88(0.08)	87(0.8)	80(0.96)	72(0.9)
	101(0)	111(0.8)/ 112(1.2)	112(0)	95(0)

^a1/2 means one molecule relative to the other one.

intermolecular vibrations fall in the same range for all acene dimers. The lower bound wavenumbers are 15–33 cm^{-1} , and the upper bound wavenumbers are 94–112 cm^{-1} . The wavenumbers for the intramolecular vibrations are found at higher wavenumbers but begin to overlap with the intermolecular vibration wavenumbers starting with anthracene where the intermolecular mode at a wavenumber of 111 cm^{-1} with an intensity of 0.8 is not readily discriminated from the intramolecular mode at 112 cm^{-1} with an intensity of 1.2, on the basis of wavenumber or intensity value. Thus, inter- and intramolecular vibrations are easily discriminated for naphthalene, and the inter- and intramolecular vibration modes are uncorrelated. For anthracene, tetracene, and pentacene, the lowest intramolecular vibration modes fall within the range of intermolecular vibrations.

Four vibration modes, given in Table 9, corresponding to rotation, translation, rotation of the dimer combined with translation and rotation movements of one molecule relative to the other, are common for all four dimers. The last two intermolecular modes are compound-specific. For tetracene, the two additional modes comprise twisting movements of molecular pairs in the same direction (87 cm^{-1}) and in opposite directions (112 cm^{-1}). For pentacene, one is a wagging movement (74 cm^{-1}) and the other comprises an opposite direction wagging movement combined with vertical translation normal to the two molecules (95 cm^{-1}). For

anthracene, one mode is a rotation of one molecule relative to the other, and the second is split into two modes. These latter two modes, at 111 and 112 cm^{-1} , consist of a butterfly movement by one molecule with a half-twisting movement by the other simultaneously.

The other vibration modes, shown in Table 10, correspond to intramolecular vibrations. For every type of vibration, simultaneous vibration of each molecule is observed, either in the same or opposite direction, resulting in two wavenumbers. These alternating motions are characterized by wavenumbers and intensities that depend on the value of the dipole moment. The butterfly vibration mode presents very different wavenumbers depending on whether the molecular movements are in the same or opposite direction. Differences range in value from 8 to 42 cm^{-1} . The impact of relative movement increases with the number of aromatic rings, due to the correlation with twisting modes at higher frequencies.

Trends in Values for Specific Bands in the Acene Family. In order to visualize the effect of the number of aromatic rings on the wavenumbers of specific vibration modes, intramolecular vibration modes and intermolecular vibrations are presented in Figures 7 and 8, respectively. Intramolecular vibration modes shift to lower wavenumbers with increasing molecule size irrespective of the vibrational mode considered, as was observed for monomers. Butterfly (B), twisting (τ), wagging (ω), stretching (σ), and rocking (ρ) movements behave similarly. By contrast, the intermolecular vibration modes, found at similar wavenumbers for all the molecules in the acene family, appear specific to this family, and may be used as a reliable fingerprint. In addition, twisting vibration bands appear for dimers of tetracene at 208 cm^{-1} and pentacene at 168 cm^{-1} that replace corresponding twisting bands for monomers at 154 and 104 cm^{-1} . The absence of a mode at 154 cm^{-1} for tetracene dimers supports the hypothesis that the feature observed experimentally at 142 cm^{-1} for monomers (see Table 3) is an overtone.

The Impact of Molecule Distortion on the Vibration Modes of Acene Dimers. The impact of molecular distortion on vibration modes for dimers was evaluated in the same manner as for monomers. The molecular conformations were

Table 10. Intramolecular Vibration Modes (cm⁻¹) and the Corresponding Intensities in Brackets (km/mol) for Naphthalene, Anthracene, Tetracene, and Pentacene Dimers^a

type ^b	naphthalene	anthracene	tetracene	pentacene
butterfly sw	188(3.8) <i>177(3.2)</i>	111(0.8)/112(1.2) <i>93(1.5)</i>	59(0.6) <i>57(0.8)</i>	42(0.4) <i>38(0.5)</i>
butterfly ow	196(0) <i>177(3.2)</i>	128(0) <i>93(1.5)</i>	94(0) <i>57(0.8)</i>	84(0) <i>38(0.5)</i>
twisting sw	207(1.2) <i>191(0)</i>	148(0.6) <i>124(0)</i>	117(0.3) <i>93(0)</i>	108(0.2) <i>74(0)</i>
twisting ow	208(0) <i>191(0)</i>	149(0.02) <i>124(0)</i>	124(0) <i>93(0)</i>	101(0.002) <i>74(0)</i>
scissoring sw	381(2.2) <i>379(1.5)</i>	244(2.1) <i>243(1.4)</i>	168(1.7) <i>168(1.3)</i>	123(1.5) <i>123(1.0)</i>
scissoring ow	380(0) <i>379(1.5)</i>	243(0) <i>243(1.4)</i>	169(0) <i>168(1.3)</i>	125(0) <i>123(1.0)</i>
wagging sw	405(0) <i>403(0)</i>	250(0) <i>276(1.1)</i>	162(0) <i>196(0)</i>	116(0) <i>150(0)</i>
wagging ow	405(0.02) <i>403(0)</i>	253(0.2) <i>276(1.1)</i>	176(0.2) <i>196(0)</i>	130(0.08) <i>150(0)</i>
twisting sw	498(0) <i>494(0)</i>	285(0.02) <i>281(0)</i>	208(0)	168(0.05)
twisting ow	499(15.2) <i>494(0)</i>	286(0.08) <i>281(0)</i>	212(0.09)	168(0.2)
wagging sw	504(45.0) <i>504(28.7)</i>	396(0.2) <i>394(0.05)</i>	282(1.6) <i>276(1.1)</i>	204(2.2) <i>198(1.3)</i>
wagging ow	507(0) <i>504(28.7)</i>	396(0.04) <i>394(0.05)</i>	286(0) <i>276(1.1)</i>	210(0.002) <i>198(1.3)</i>
stretching sw	534(0) <i>534(0)</i>	407(0) <i>407(0)</i>	324(0) <i>324(0)</i>	269(0) <i>269(0)</i>
stretching ow	534(0.09) <i>534(0)</i>	407(0.04) <i>407(0)</i>	325(0.06) <i>324(0)</i>	269(0.05) <i>269(0)</i>
rocking sw	537(0) <i>537(0)</i>	412(0) <i>412(0)</i>	316(0) <i>316(0)</i>	246(0) <i>247(0)</i>
rocking ow	537(0.06) <i>537(0)</i>	412(0.01) <i>412(0)</i>	316(0.03) <i>316(0)</i>	247(0.05) <i>247(0)</i>
twisting sw	658(0.4) <i>656(0)</i>		335(0.03) <i>327(0)</i>	256(0.02) <i>247(0)</i>
twisting ow	659(0) <i>656(0)</i>		336(0) <i>327(0)</i>	257(0.008) <i>247(0)</i>
wagging sw		496(70.5) <i>495(0)</i>	390(0) <i>389(0)</i>	302(0.005) <i>247(0)</i>
wagging ow		497(0.5) <i>495(0)</i>	390(0.02) <i>389(0)</i>	303(0.07) <i>247(0)</i>

^aValues of corresponding monomer modes are written below in italics. ^bsw = same way (i.e., resulting dipole moment in the same direction for the two molecules). ow = opposite way (i.e., resulting dipole moment in opposite directions for the molecules).

perturbed as illustrated in Figure 3, and the distance r (defined in Figure 6) between the molecules was set equal to 3.4, 3.4, and 2.9 Å for naphthalene, anthracene, and tetracene dimers, respectively. As distortion reduces the aromaticity of the molecules, the stability of π -stacked dimers is expected to be reduced or eliminated. Table 11, where interaction energies for these systems are listed, highlights this loss of stability. For example, no conformations for distorted pentacene dimers correspond to local minima in the PES. Distortion did not affect intramolecular modes of vibration overall for dimers, as shown in Table S3 in the Supporting Information. Only butterfly vibration modes for anthracene and tetracene, whose wavenumbers get closer to the ones observed for monomers, shown in Table 10, appear to be affected. Some other modes, for which two wavenumbers are given in Table S3 (Supporting Information), do not exhibit the combined movement of the two molecules either in the same direction or in the opposite direction (as observed for the other intramolecular modes) but movement of the molecules simultaneously. These exceptions are a consequence of the decoupling between the molecules forming the distorted dimer. However, the intermolecular vibration modes, shown in Table 12, are shifted to lower

wavenumbers compared to the symmetric analogues shown in Table 9. For the symmetric dimers, the intermolecular vibration modes are effectively independent of molecule size, while, for the distorted molecules, the vibration modes tend to lower values as molecular size increases.

The behavior of the symmetric dimers is attributed to a balancing of opposing effects. One would expect intermolecular vibration modes to shift to lower wavenumbers with increasing ring number (as is observed for intramolecular modes). However, the intensity of the intermolecular interactions increases with the number of rings (Table 8) for symmetric dimers, and this tends to increase wavenumbers. This compensative effect is diminished or absent for the distorted dimers; hence, the difference in behavior and wavenumber invariance for a given intermolecular vibration mode may be a characteristic of π -stacking interactions.

Identification of Unassigned Peaks in the Experimental Spectra of Tetracene and Pentacene. By subtracting the intermolecular wavenumbers from the symmetric dimer vibrational spectra of tetracene and pentacene, and comparing the resultant spectra with the spectra of the corresponding monomers, there are two peaks, one for tetracene at 117 cm⁻¹

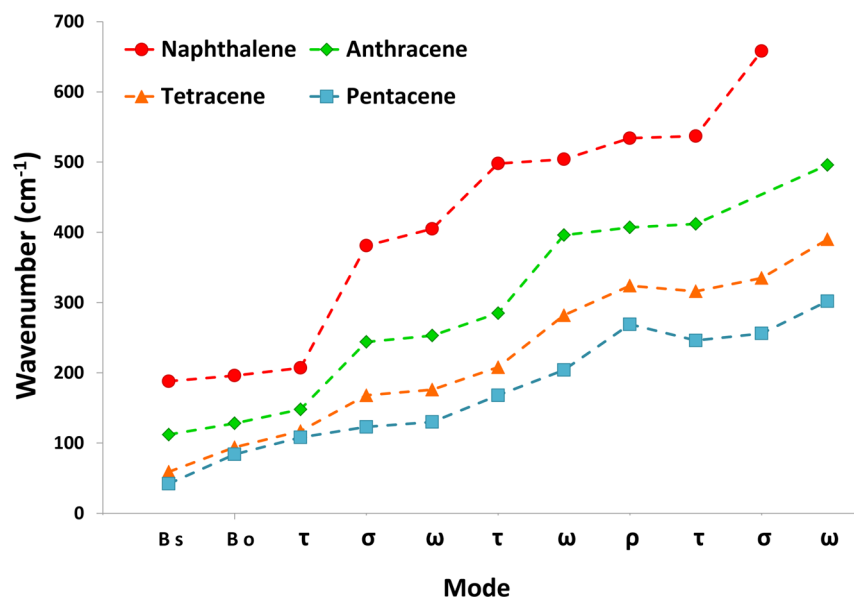


Figure 7. Comparison of theoretical intramolecular vibration modes for naphthalene, anthracene, tetracene, and pentacene dimers: B s and B o designate butterfly modes, same (s) and opposite (o) direction, respectively; τ , σ , ω , and ρ take place for twisting, stretching, wagging, and rocking modes, respectively.

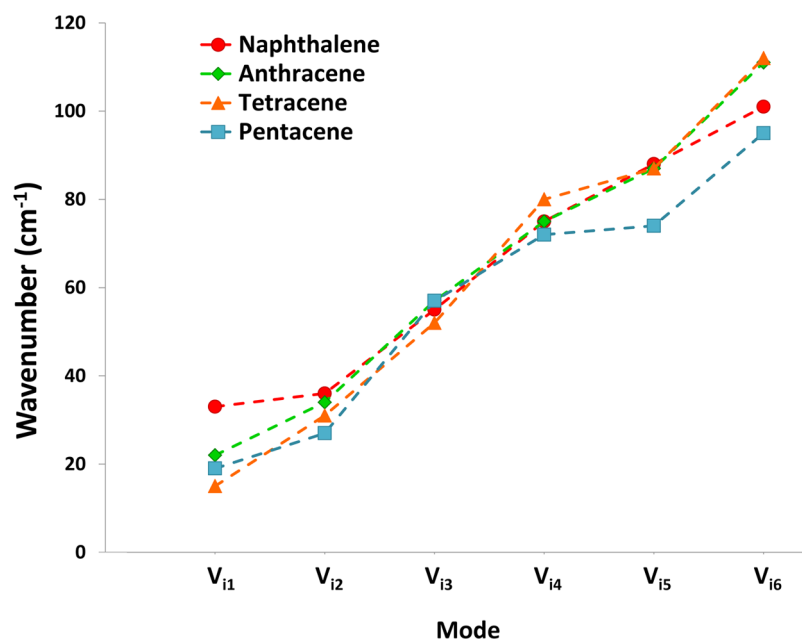


Figure 8. Comparison of theoretical intermolecular vibration modes for naphthalene, anthracene, tetracene, and pentacene dimers.

Table 11. Interaction Energies Calculated for Distorted Acene Dimers^a

dimers	naphthalene	anthracene	tetracene
interaction energies (kcal/mol)	+1.23 (+8.76)	-4.27 (+7.25)	-10.81 (+4.92)

^aThe values in parentheses indicate the energy difference between the interaction energy of distorted and symmetric acene dimers.

and one for pentacene at 108 cm^{-1} , with intensities in the same order of magnitude as the ones calculated for fundamental bands for tetracene (0.34 km/mol) and pentacene (0.15 km/mol). These peaks arise because, for every vibrational mode occurring for monomers, there are two quasi-degenerate modes arising for dimers depending on whether the molecules vibrate in the same or opposite direction. Thus, modes that are Raman-

active for a monomer become infrared-active for dimers vibrating in opposite directions.

With this final set of assignments, the intermolecular and intramolecular vibration analysis for the acene family of compounds is complete. Figure 9 provides a direct comparison of experimental and calculated spectra with the finalized assignments. There are no residual unassigned peaks in the available experimental spectra, and ultra-low-wavenumber

Table 12. Calculated Intermolecular Vibrational Modes (cm^{-1}) and Their Intensities for Distorted Dimers of Naphthalene, Anthracene, and Tetracene

type ^a	naphthalene	anthracene	tetracene
translation 1/2	12(0.015)	11(0.004)	6(0)
rotation 1/2	32(0.04)	21(0.004)	13(0.004)
translation 1/2	47(0.007)	33(0.002)	21(0.005)
rotation + translation	60(0.006)	41(0.009)	33(0.003)
translation (remoteness of molecules)	73(0.008)	55(0.01)	46(0.2)
rotation 1/2	94(0.14)	73(0.2)	
twisting of the full system			74(0.08)

^a1/2 means one molecule relative to the other.

intermolecular peaks are hypothesized. While these have yet to be measured experimentally, identification and discrimination of intermolecular and intramolecular bands at low wavenumbers will play a more important role in the future as the range of IR measurements broadens. For example, by using coherent synchrotron radiation, it is becoming possible to make experimental photoacoustic infrared measurements in the range $7\text{--}30\text{ cm}^{-1}$, as demonstrated recently by Billingham and Michaelian.¹⁰⁷ DFT calculations have much to contribute to this development, and the interpretation of results obtained.

CONCLUSIONS

The DFT calculations presented in this work provide detailed and accurate descriptions of experimental photoacoustic far-infrared spectra for tetracene and pentacene and for far-infrared spectra of the acene family as a whole. The impacts of molecular distortion on monomer and dimer spectra and dimer stability and the interpretation of peaks in photoacoustic

infrared spectra of crystalline solids are discussed in detail. Distortion breaks molecular symmetry, and this adds to the complexity of the interpretation of the resulting spectra. Intermolecular and intramolecular vibrations in the acene family of compounds are discriminated, and combinations of intermolecular and intramolecular vibrations are identified. Peak assignment ambiguity and the potential impacts of molecular distortion on peak intensity are addressed for tetracene and pentacene. By examining computed and experimental spectra for the family of acene molecules concurrently, trends in vibration modes with molecular size were used to reinforce assignments, and to reattribute assignments made previously in the experimental tetracene and pentacene spectra. In particular, dimer calculations permitted reassignment of an infrared-active peak arising in tetracene at 142 cm^{-1} as an overtone, and assignment of large infrared-active peaks at 117 and 108 cm^{-1} for tetracene and pentacene, respectively, that may permit their identification in an acene mixture.

The intermolecular vibrations for naphthalene, anthracene, tetracene, and pentacene are found in the same wavelength range of the far-infrared spectrum and are not readily discriminated one from the other. The specificity of the acene π -stacking interactions, reinforced by calculations with distorted dimers, not only allows the identification of this family of molecules but also highlights the presence of π -stacking interaction in a mixture as a whole.

ASSOCIATED CONTENT

Supporting Information

Detailed aspects of tetracene (molecule) geometry and computed outcomes referred to in the text as Tables S1–S3

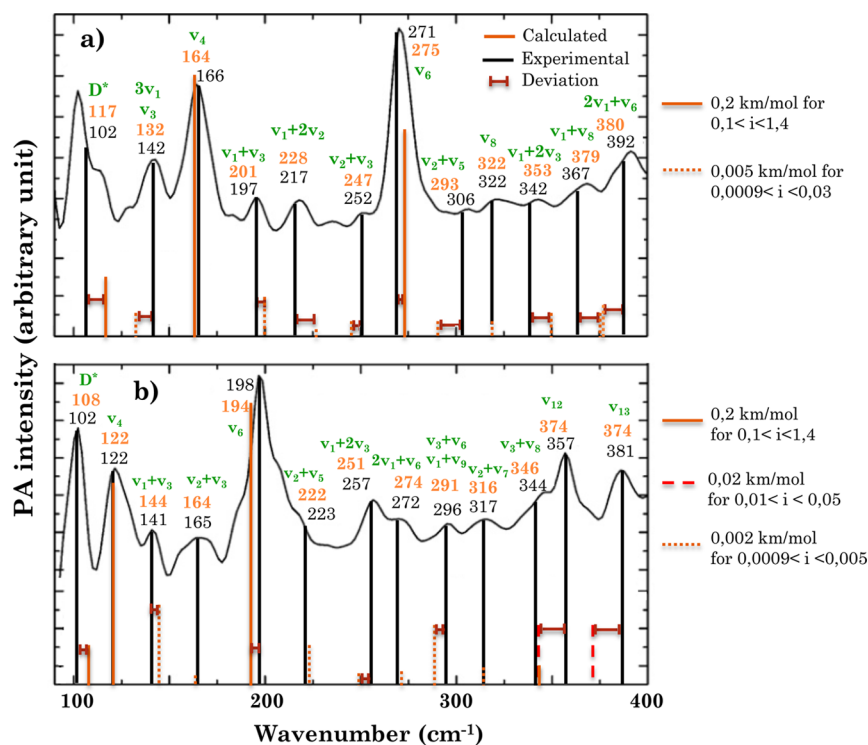


Figure 9. Final assignments of the experimental infrared transitions for (a) tetracene and (b) pentacene. Calculated wavenumbers are indicated on the spectra with orange lines, with corresponding values above. Experimental values are indicated for comparison, and represented with dark lines. Computed intensities are scaled in order to provide qualitative information.

and Figures S1 and S2. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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Notes

The authors declare no competing financial interest.

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